

## Unit 2- Ch 9 Naming Chemicals

Chem111/112

\*\*How do you name? Make up a page with an example for each category.\*\*

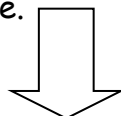
1.Binary Ionic? 2.Ternary Regular? 3.Ternary Irregular?	Hydrates?	Bases	1.Binary Acids 2.Ternary Regular Acids 3.Ternary Irregular?	Molecular 1. Diatomics 2. Binary 3. Polyatomics 4. List ones that you had to study ex. H <sub>2</sub> O <sub>2</sub>
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Give some diagnostic test used to distinguish between acids and Bases.

Fill in the missing state or formula or name.

Fomula(state)	Name	Formula(state)	Name:
1.	Sodium sulfate decahydrate	18 Fe <sub>2</sub> O <sub>3</sub> (s)	
2 NH <sub>4</sub> NO <sub>3</sub> (aq)		19	nitrogen
3	phosphorous trihydride	20 SO <sub>3</sub> (g)	
4 H <sub>3</sub> PO <sub>5</sub> (aq)		21	ammonium chloride
5	silicon dioxide	22 P <sub>4</sub> (s)	
6 HF(aq)		23	carbon tetrachloride
7 N <sub>2</sub> H <sub>4</sub> (g)		24 C <sub>2</sub> H <sub>5</sub> OH(l)	
8	perchloric acid	25	Methane
9	nickel(II) bromide	26	acetic acid
10 PbO <sub>2</sub> (s)		27	nitrous acid
11	tetra phosphorous decaoxide	28 HCl(aq)	
12 KNO <sub>3</sub> (aq)		29 HClO <sub>4</sub> (aq)	
13	sucrose	30 KCl(aq)	
14	propane	31	calcium carbonate
15 SiO <sub>2</sub> (s)		32 C(s)	
16 CaSO <sub>2</sub> •2H <sub>2</sub> O(s)		33	Iodine
17	Lead(IV)oxide	34 FeSO <sub>4</sub> (s)	

→Using a different colour pen write i-ionic, a-acid, b-base, m-molecular in front of each formula above. More on next page.



→ Attach a page folded in four with the properties, diagnostic tests, states and how to recognize each category of chemical from its formula. Write all of the formulas from above into your table.

Ionic Properties	Ex.FeSO <sub>4</sub> (s) list others	
Diagnostic tests	from above	
State		
How to recognize?		